

Chapter 7 Review Modern Chemistry Answers

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Modern Chemistry • CHAPTER 7 HOMEWORK 7-2 ... CHAPTER 7 • HOMEWORK 7-2 VOCABULARY 1. d 2. c 3. e 4. a 5. g 6. h 7. b 8. f GRAPHIC ORGANIZER Name of Compound or Ion Formula of Compound Copper(II) ion Cu^{2+} Ammonium ion NH_4^+ Bismuth(III) bromide BiBr_3 Stannous fluoride SnF_2 Magnesium fluoride MgF_2 Iron(III) cyanide $\text{Fe}(\text{CN})_3$

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HOMEWORK 7-2

CHAPTER 6 REVIEW Chemical Bonding MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. a.

7 Chemical Formulas and Chemical Compounds

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO_3 a. copper(II) carbonate Na_2SO_3 b. sodium sulfite $(\text{NH}_4)_3\text{PO}_4$ c. ammonium phosphate SnS_2 d. tin(IV) sulfide HNO_2 e. nitrous acid 2.

5 The Periodic Law

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6 Chemical Bonding

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Title: Modern Chemistry Chapter 7 Chemical Formulas 1 Modern Chemistry Chapter 7 Chemical Formulas Chemical Compounds . A chemical formula indicates the kind and relative number of atoms in a chemical compound. C_8H_{18} (octane) has 8 carbon and 18 hydrogen atoms. 2 Forming Ionic Compounds. Compounds that have the elements held together by

7 Chemical Formulas and Chemical Compounds

Modern Chemistry 57 Chemical Bonding CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: ____ a. If the formula mass of one molecule is x u, the molar mass is x g/mol. ____ b.

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on ...

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PPT - Modern Chemistry Chapter 7 Chemical Formulas ...

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered.