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# Chapter 5 Review The Periodic Law

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## CHAPTER 5 REVIEW

### The Periodic Law

#### SECTION 5-3 SHORT

ANSWER Answer the following questions in the space provided. 1.

When an electron is added to a neutral atom, energy is . (a) always absorbed (c) either absorbed or released (b) always released (d) burned away

2. The energy required to remove an electron from an atom is the atom's .

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The periodic law states

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that the properties of elements are periodic functions of their atomic numbers. This means that the \_\_\_\_ determines the \_\_\_\_ position of each element in the periodic table. This means that the \_\_\_\_ determines the position of each element in the periodic table.

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Modern Chemistry 5  
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CHAPTER 5 REVIEW  
The Periodic Law  
SECTION 2 SHORT  
ANSWER Use this  
periodic table to  
answer the following  
questions in the space  
provided. 1. Identify  
the element and write  
the noble-gas notation  
for each of the  
following: a. the Group  
14 element in Period 4

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the elements in a chart  
Organized about 70  
elements by increasing  
atomic mass Left blank  
spaces for elements  
which were not  
discovered yet .

Mendeleev's Periodic  
Table. Periodic Table  
Henry Moseley (1911) -  
rearranged  
Mendeleev's periodic  
table to the current ...

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CHAPTER 5 REVIEW  
THE PERIODIC LAW

SECTION 1 1. \_\_\_\_\_ In the modern periodic table, elements are ordered (a) according

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to decreasing atomic mass. (b) according to Mendeleev's original design.

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The left side of the periodic table is composed of metals, which have few valence electrons. Moving to the right, there are more electrons and their mobility is less, reducing their metallic properties. When most of the orbitals are filled, the atoms take on nonmetallic characteristics. When orbitals are full, a noble gas configuration is

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attained. Metals  
become positive ions.  
When sodium ...

## **Answers to Topic 5 Periodic Table from Review Book**

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CHAPTER 5 REVIEW

The Periodic Law

SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1.

c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original

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design. (c) according to increasing atomic number. (d) based on when they were discovered.

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Holt McDougal Modern  
Chemistry 1 Chapter  
Test Assessment

Chapter Test A

Chapter: The Periodic  
Law Use the periodic  
table below to answer

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the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

\_\_\_\_\_ 1. Mendeleev organized the chemical

...

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SECTION 2 SHORT  
ANSWER Use this  
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answer the following  
questions in the space  
provided. 1. Identify  
the element and write  
the noble-gas notation  
for each of the  
following: a. the Group  
14 element in Period 4

## **Chapter 5 Review The Periodic Law Answers 5 1**

CHAPTER 5 REVIEW \

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## SHORT ANSWER

Answer the following questions in the space provided. 1. Consider the neutral atom with 53 protons and 74 neutrons to answer the following questions. 53. a. What is its atomic number? 127. b. What is its mass number? atomic number. c. Is the element's position in a modern periodic table ...

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